ADDITIONS AND CORRECTIONS

Transition State Differences in Hydrolysis Reactions of Alkyl versus Aryl Phosphate Monoester Monoanions [*J. Am. Chem. Soc.* **2003**, *125*, 13106−13111]. Piotr K. Grzyska, Przemyslaw G. Czyryca, Jamie Purcell, and Alvan C. Hengge*

Page 13107, Figure 3. The reported isotope effect $^{18}k_{\text{bridge}}$ for hydrolysis of the monoanion of *m*NBP is in error; the correct value is 1.0157 ± 0.0009. For comparison, $^{18}k_{\text{bridge}}$ for *p*NPP monoanion hydrolysis is 1.0087. Since the maximum $^{18}k_{\text{bridge}}$ for the alkyl leaving group *m*NBP is expected to be about twice[1] that for the aryl leaving group in *p*NPP, the KIEs for the two monoanion reactions are similar in proportion to their respective upper limits. This indicates a similar extent of P−O bond fission in the two reactions. The other data reported in the paper are unaffected. The major conclusion, that the data are most consistent with a pre-equilibrium proton transfer followed by rate-determining P−O bond fission for both *p*NPP and *m*NBP, is unchanged.

References

(1) Sawyer, C. B.; Kirsch, J. F. *J. Am. Chem. Soc.* **1973**, *95*, 7375−7381. Gorenstein, D. G.; Lee, Y. G.; Kar, D. *J. Am. Chem. Soc.* **1977**, *99*, 2264−2267. O'Leary, M. H.; Marlier, J. F. *J. Am. Chem. Soc.* **1979**, *101*, 3300−3306.

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Rates of Oxygen Exchange between the $[H_xNb_6O_{19}]^{8-x}{}_{(aq)}$ Lindqvist Ion and Aqueous Solutions [*J. Am. Chem. Soc.* **2006**, *128*, 14712−14720]. Jay R. Black, May Nyman, and William H. Casey*

Page 14716, Table 2. Due to an error in data reduction, the activation parameters in Table 2 were incorrect. The correct values are given in the table below.

Table 2. Activation Enthalpies and Entropies for Oxygen Exchange from Water with the η-O and μ_2-O(H) sites of the Hexaniobate Ion as a Function of pH

	η-O		μ_2-O(H)	
pH	$\Delta H^\ddagger \pm 2\sigma$ (kJ mol^{-1})	$\Delta S^\ddagger \pm 2\sigma$ (J mol^{-1} K^{-1})	$\Delta H^\ddagger \pm 2\sigma$ (kJ mol^{-1})	$\Delta S^\ddagger \pm 2\sigma$ (J mol^{-1} K^{-1})
∼13	89.4 ± 4.8	−42.9 ± 16.1	88.0 ± 2.8	−29.7 ± 9.4
∼14.5	93.9 ± 2.7	−25.8 ± 9.0	104.2 ± 7.6	10.2 ± 25.4

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Vanadium-Catalyzed Asymmetric Epoxidation of Homoallylic Alcohols [*J. Am. Chem. Soc.* **2007**, *129*, 286−287]. Wei Zhang and Hisashi Yamamoto*

The Supporting Information published with the Communication was incomplete. The missing portion has now been made available.

Supporting Information Available: Representative experimental procedures and spectral data for **1c** and **1d**. This material is available free of charge via the Internet at http://pubs.acs.org.

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Reactivation of the p53 Tumor Suppressor Pathway by a Stapled p53 Peptide [*J. Am. Chem. Soc.* **2007**, *129*, 2456−2457]. Federico Bernal, Andrew F. Tyler, Stanley J. Korsmeyer, Loren D. Walensky,* and Gregory L. Verdine*

Page 2456, Figure 1. The correct sequence of the peptide UAH-p53-**8** in Figure 1B should read

Ac-QSQQTF*NLWRLL*QN-NH_2

which is identical to the sequence of peptide SAH-p53-**8**. The only difference between these two compounds is that UAH-p53-**8** is "unstapled".

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